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17. Without the salt bridge, the circuit would be open (or broken)

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and no current could flow. With a salt bridge, each half-cell remains electrically neutral and current can flow through the circuit. 19.

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1. The instantaneous rate is the rate of a reaction at any particular point in time, a period of time that is so short that the concentrations of reactants and products change by a negligible amount.

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Chapter 17. Additional Aspects of Equilibrium. 17.1 The Common Ion Effect. • The dissociation of a weak electrolyte is decreased by the addition of a strong electrolyte that has an ion in common with the weak electrolyte. • For example, consider the ionization of a weak acid, acetic acid. HC. 2H.

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This is the summary for Chapter 17 "Additional Aspects of Aqueous Equilibria" of Brown et al. Textmap, which addresses more complex equilibria with solutions containing more than one

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solute.

17: Additional Aspects of Aqueous Equilibria - Chemistry

...

Chapter 17. Electrochemistry. 17.6 Corrosion. Learning Objectives. By the end of this section, you will be able to: ...
Answers to Chemistry End of Chapter Exercises. 2. Mg and Zn. 4. Both examples involve cathodic protection. The (sacrificial) anode is the metal that corrodes (oxidizes or reacts). In the case of iron (-0.447 V) and zinc ...

17.6 Corrosion - Chemistry

This is the lecture recording for Chapter 17 in John McMurry's Organic Chemistry: Alcohols

Organic Chemistry - Chapter 17 - McMurry - Alcohols - YouTube

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This chapter introduces you to thermochemistry, a branch of chemistry that describes the energy changes that occur during chemical reactions. In some situations, the energy produced by chemical reactions is actually of greater interest to chemists than the material products of the reaction.

5: Thermochemistry - Chemistry LibreTexts

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