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Chemistry Packet Answer Key

# Unit 9 Solutions Chemistry Packet Answer Key

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## **Unit 9 Solutions Chemistry Packet**

1. Calculate the molarity of each of the following solutions: (a) 2.5 mol of NaOH in 0.500 L of solution (b) 1.8L of solution

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containing 3.3 mol  $\text{KNO}_3$  (c) 3.0 mol of  $\text{NaOH}$  in 0.500 L of solution (d) 12 g of  $\text{HCl}$  in 250 mL of solution 2. Calculate the total moles of solute in each of the following solutions:

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(a) 2.5 mol of  $\text{NaOH}$  in 0.500 L of

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solution (b) 1.8L of solution containing  
3.3 mol  $\text{KNO}_3$  (c) 30. g of NaOH in 0.500  
L of solution (d) 522 g of  $\text{K}_2\text{SO}_4$  in 1.5  
L of solution (e) 12 g of HCl in 250 mL of  
solution 2. Calculate the total moles of  
solute in each of the following solutions:  
2 Calculate the concentration of various  
solutions

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## **Practice Packet Unit 9: Solutions**

(1) an element and a mixture (3) a solution and a mixture (2) an element and a compound (4) a solution and a compound 12. Two different samples decompose when heated. Only one of the samples is soluble in water. Based on this information, these two samples are (1) both the same element (3) both

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the same compound

## **Practice Packet Unit 9: Solutions**

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Acids, Bases, & Salts \*KEY\* \* KEY \* 2  
Unit Vocabulary: Amphoteric Arrhenius  
acid ... solution Understand how pH  
works Using Table M, determine the pH  
of a given solution Unit 9: Acids, Bases,  
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## **Unit 9 Packet and Key - Ms. Kovach's Chemistry Class**

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Unit 9: Solutions  
Review Molarity Calculations 1) 2.3

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moles of sodium chloride in 0.45 liters of solution. 2) 1.2 moles of calcium carbonate in 1.22 liters of solution. 3) 98 grams of NaOH in 2.2 liters of solution. 4) 1.2 grams HCl in 25 mL of solution.

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3 Review-ANSWERS

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Vocabulary: Alloy Aqueous boiling point  
boiling point elevation colligative  
properties colloid concentrated  
concentration dilute freezing point  
depression heterogeneous  
homogeneous insoluble miscible mixture  
molarity (M) parts per million (ppm)  
percent by mass percent by volume  
precipitate saturated solubility

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preferably in a small group where you can draw from each other's understanding.

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